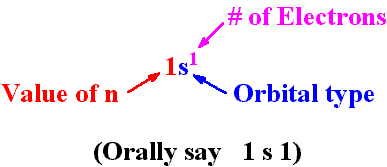
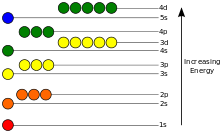
Orbitals:

Orbital-

|  |  |
| --- | --- |
| Types of orbits | How many |
| s |  |
| p |  |
| d |  |
| f |  |
| g |  |

|  |  |  |
| --- | --- | --- |
| Energy Level | Types of Orbitals | Orbitals |
|  |  |  |
|  |  |  |
|  |  |  |
|  |  |  |
|  |  |  |



Rule #1 fill the available orbital with lowest energy first (Aufbau Principle)

Rule #2 a maximum of two electrons can occupy any orbital. ( Paul-Exclsion Principlea0

Rule #3 If more than one orbital have the same energy, fill an empty one first (Hund’s rule)

|  |  |
| --- | --- |
| Oxygen (\_\_\_\_\_\_\_electrons)  Orbital representation diagram.svg | Argon (\_\_\_\_ electrons)  Orbital representation diagram.svg |
| Arsenic (\_\_\_\_electrons)  Orbital representation diagram.svg | Calcium ion (\_\_\_\_electrons)  Orbital representation diagram.svg |